Naming molecular compounds worksheet with answers

When you name ionic compounds, you write the name of the metal first and then the nonmetal. Suppose that you want to name the compound that results from the reaction of lithium and sulfur. You first write the name of the metal, lithium, and then write the name of the nonmetal, adding an -ide ending so that sulfur becomes sulfide. Ionic compounds involving polyatomic ions follow the same basic rule: Write the name of the metal first, and then simply add the name of the nonmetal (with the polyatomic anions, it is not necessary to add the -ide ending). When the metal involved is a transition metal with more than one oxidation state, there can be more than one way to correctly name the compound. For example, suppose that you want to name the compound formed between the cation: and the cyanide ion: The preferred method is to use the metal name followed in parentheses by the ionic charge written as a Roman numeral: Iron(III). But an older naming method, which is still in use, is to use -ous and -ic endings. The ion with the lower oxidation state (lower numerical charge, ignoring the + or -) is given an -ous ending, and the ion with the higher oxidation state (higher numerical charge) is given an -ic ending. So the compound can be named: Sometimes figuring out the charge on an ion can be a little challenging (and fun), so suppose you want to name the following ionic compound: The sulfate ion has a 2- charge, and from the formula you can see that there are two of them. Therefore, you have a total of four negative charges. The ammonium ion has a 1+ charge, so you can figure out the charge on the iron cation. Because you have a 4- for the sulfates and a 1+ for the ammonium, the iron must be a 3+ to make the compound neutral. So the iron is in the Iron(III), or ferric, oxidation state. You can name the compound: And, finally, if you have the name, you can derive the formula and the charge on the ions. For example, suppose that you're given the name cuprous oxide. You know that the cuprous ion is: The oxide ion is: Applying the crisscross rule, you get the following formula: The Social Security Administration (SSA) compiles a list of the most popular baby names over the past 100 years. This represents perhaps the most complete picture of the most common names in the United States. The following list includes both male and female names — ranked from 10 to 1— as well as the most common last names based on the 2010 census.Charles and Margaret MartinezIn 2010, 1,060,159 people in the US had the last name Martinez. This was the tenth most common last name in the country. Meanwhile, the tenth most common first names over the last 100 years have been Charles (2,144,937 or 1.23 percent of 173,916,919 male births) and Margaret (993,136 or 0.59 percent of 169,671,039 female births).Thomas and Sarah RodriguezThe ninth most common last name was also Hispanic. There were 1,094,924 Rodriguezes in 2010. And of all the baby boys born between 1918 and 2017, 2,174,023 have been Thomases. A total of 996,554 baby girls were named Sarah.Joseph and Jessica DavisTwo common 'J' names are the eighth most popular in America: the biblical Joseph and Jessica. These aCC0unted for 2,384,205 and 1,043,436 baby names, respectively. Meanwhile, in 2010, 1,116,357 people had the last name Davis.Richard and Susan MillerA sum of 2,487,983 Richards amount to 1.43 percent of all baby boys born between 1918 and 2017. The 1,106,071 Susans represent 0.65 percent of the total baby girl births. As for Millers, there were 1,161,437 of these recorded in 2010.David and Barbara GarciaAnother Hispanic last name comes in at number six. There were 1,166,120 Garcias in the 2010 census. But over the past century, there have been more than 3,557,293 Davids and over 1,410,059 Barbaras.William and Elizabeth JonesLove them or loathe them, the British royal family have long been the inspiration for baby names. There were 3,662,399 Williams born between 1918 and 2017 and 1,443,415 Elizabeths. Good luck keeping up with the Joneses, though; there were 1,425,470 of these in 2010.Michael and Linda BrownThe fourth most common last name in 2010 was Brown, with 1,437,026 occurrences. Michael and Linda were the fourth most common baby names in the century before 2018. There were 4,315,462 and 1,448,097 of these, respectively.Robert and Jennifer WilliamsRobert and Jennifer are both pretty versatile names. Diminutives include Rob, Bob, Robbie, Bobby, and Jen, Jenna, Jenny and Jennie. That might be why they're so popular. Of all baby boys born 1918-2017, 4,571,203 (2.63 percent) were called Robert, while 1,465,928 (0.86 percent) of all girls were called Jennifer. 1,625,252 people had the last name Williams in 2010.John and Patricia JohnsonOK, so you might not find many John Johnsons; but they both take the number two spot. A total 2.64 percent of baby boys (4,594,023 born in the last century) were given the name John, and the last census recorded 1,932,812 last names as Johnson. The 1,564,163 Patricias (Pat, Patty, Trisha, Trixie) represent 0.92 percent of all female births between 1918 and 2017.James and Mary SmithNo surprises here: Smith is by far the most common name in the USA. In the 2010 census, 2,442,977 of them were counted. James and Mary are the most common first names. But while there were just 198,931 more Jameses than Johns born 1918-2017, there were 1,829,293 more Marys than Patricias. The 3,393,456 females called Mary aCC0unted for 2 percent of all girl births. MORE FROM QUESTIONSANSWERED.NET Jarmoluk/Pixabay While the compound P4O10 has many names, its most common name is phosphorus pentoxide. This compound is made up of four phosphorus atoms and 10 oxygen atoms bonded together with covalent bonds.How Was P4O10 Named? The name may be confusing to some, since the prefix "pent-" refers to five, and there are 10 oxygen atoms in the compound. Phosphorus pentoxide gets its name from the empirical formula of the compound, which is P2O5. Molecules of P2O5 are unstable and associate with each other to form the larger molecules of P4O10. Physical Properties of P4O10 At room temperature, phosphorus pentoxide is a solid, white waxy substance and is typically a powder. The molecule has a hexagonal shape and is held together with weak van der Waals forces. This compound is unique in that it exists in four different polymorphs. The most common is two molecules of P2O5 joined together to form the larger P4O10. All of phosphorus pentoxide's polymorphs are based around the tetrahedral arrangement of the phosphorus and oxygen atoms make up the compound. It melts at 340C and boils at 360C. Because the melting point and boiling points are so close together, phosphorus pentoxide often skips melting and sublimates from solid to gas. Chemical Properties of P4O10 Phosphorus pentoxide is a polar compound. It is a noncombustible compound, meaning that it does not react with oxygen to produce a flame. However, it is highly reactive with water and forms phosphoric acid when combined with water. It can cause a fire if it comes in contact with water-containing materials, such as cotton or wood. When exposed to metal, it causes corrosion and forms various metal oxides. It's also corrosive to skin and other tissues, leading to chemical burns and respiratory inflammation. These irritations and injuries occur even in small concentrations. Proper safety measures are required when handling phosphorus pentoxide. Production of P4O10 Phosphorus pentoxide is typically produced through the combustion of phosphorus and oxygen. Burning tetraphosphorus with a large amount of oxygen produces the compound. White phosphorus is necessary as the starting material. Although adding water to phosphorus pentoxide produces phosphoric acid, the reaction does not work in reverse. It's not possible to produce phosphorus pentoxide by dehydrating phosphoric acid. Uses for P4O10 Phosphorus pentoxide is commonly used in the production of phosphoric acid. Because it reacts so easily with water, it's also used as a desiccant and dehydrating agent. It draws moisture out of the air to keep an area free from moisture. Phosphorus pentoxide has a tendency to form a protective layer around the outside during storage, which prevents it from drawing in any more moisture. Because of this, it's often used in its granular form for drying. As an industrial desiccant, this compound plays an intermediate role in turning acids into their anhydride counterparts. For example, it's used to convert nitric acid (HNO3) into nitrogen pentoxide (N2O5). Phosphorus pentoxide is also used in the manufacture of glass, rubber, and some laboratory procedures. Other Names for P4O10 While phosphorus pentoxide is the most common name for P4O10, it does go by other names, including: Diphosphorus pentoxide Phosphorus(V) oxide Phosphoric anhydride Tetraphosphorus decaoxide Tetraphosphorus decoxide MORE FROM REFERENCE.COM Thank you for your participation!

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